Answers  Problem Set
Nov 11, 2011

\[ \text{Na}^+ \text{O\text{NH}_2} \]

\[ \text{yields } 3 \text{ resonance forms} \]

\[ \text{yields } 2 \text{ resonance forms} \]

\[ \text{Most resonance stability} \]

\[ + \text{NH}_3 \]

\[ \text{Most stable base is weakest base} \]

\[ \text{Strong base} \rightarrow \text{Weaker base} \]
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[Chemical structures and reactions]

Strong conc. base

Most acidic hydrogen
Why leads to most stable base

Weaker base charge distributed

+ H2O
Most Resonance Stability

Charge on 3 oxygens
Which is stronger acid

Draw Conjugate base

6 resonance forms

Weaken base

More common

4 resonance forms

Rest of key coming tomorrow